

Student Name: _____ Pd. _____ Date: _____

Comprehensive Nomenclature Practice 2

The following exercises include a mix of nomenclature exercises, including binary, polyatomic, multivalent, acids, and molecular compounds. When the name of a compound is given, write the formula. When a formula is given, write the name of the compound. If the compound contains a multivalent metal, write both the Stock and Classical name when appropriate.

1. Strontium iodide
2. Iron(II) oxide
3. Copper(I) nitrate
4. Ni_2O_3
5. Sodium dihydrogen phosphate
6. Barium chloride
7. Calcium acetate
8. PF_5
9. Ammonium oxalate
10. MgCO_3
11. BaO_2
12. Potassium chlorate
13. Lead(II) hydroxide
14. Hypoiodous acid
15. Mercury(I) nitrate
16. Calcium carbide
17. Potassium sulfide
18. Bromine monochloride
19. Tin(II) carbonate
20. Au_2O_3
21. Aluminum chromate
22. Plumbous dichromate
23. Dichromic acid
24. Carbonic acid
25. $\text{Fe}(\text{NO}_2)_3$
26. Chromic acid
27. CaF_2
28. P_4O_{10}
29. K_2S
30. H_2
31. NaHCO_3
32. Aluminum selenide
33. N_2O
34. HClO
35. O_2F
36. SBr_6
37. Li_2Te
38. Carbon monoxide
39. Carbon dioxide
40. P_2S_3
41. Silver chloride
42. Zinc nitrate
43. Sodium nitride
44. Xenon hexafluoride
45. I_2
46. Ammonium nitrate
47. Tin(II) oxide
48. Sodium hydroxide
49. NO_2
50. BF_3

Comprehensive Nomenclature Practice 2 SOLUTIONS

51. Strontium iodide
Binary, SrI₂
52. Iron(II) oxide
Multivalent, FeO
53. Copper(I) nitrate
Multivalent/polyatomic, CuNO₃
54. Ni₂O₃
Multivalent, Nickel(III) oxide
55. Sodium dihydrogen phosphate
Polyatomic, NaH₂PO₄
56. Barium chloride
Binary, BaCl₂
57. Calcium acetate
Polyatomic, Ca(C₂H₃O₂)₂
58. PF₅
Molecular, phosphorus pentafluoride
59. Ammonium oxalate
Polyatomic, (NH₄)₂C₂O₄
60. MgCO₃
Polyatomic, magnesium carbonate
61. BaO
Binary, barium oxide
62. Potassium chlorate
Polyatomic, KClO₃
63. Lead(II) hydroxide
Multivalent/polyatomic, Pb(OH)₂
64. Hypoiodous acid
Ternary Acid, HIO
65. Mercury(I) nitrate
Multivalent/Polyatomic, HgNO₃
66. Calcium carbide
Binary, Ca₂C (C is 4- in this case)
67. Potassium sulfide
Binary, K₂S
68. Bromine monochloride
Molecular, BrCl
69. Tin(II) carbonate
Multivalent/Polyatomic, SnCO₃
70. Au₂O₃
Multivalent, gold(III) oxide
71. Aluminum chromate
Polyatomic, Al₂(CrO₄)₃
72. Plumbous dichromate
Multivalent/Polyatomic, PbCr₂O₇
73. Dichromic acid
Ternary Acid, H₂Cr₂O₇
74. Carbonic acid
Ternary Acid, H₂CO₃
75. Fe(NO₂)₃
Multivalent/Polyatomic, iron(III) nitrate
76. Chromic acid
Ternary Acid, H₂Cr₂O₄
77. CaF₂
Binary, calcium fluoride
78. P₄O₁₀
Molecular, tetraphosphorus decoxide
79. K₂S
Binary, potassium sulfide
80. H₂
Molecular, hydrogen (just hydrogen, see the POGIL on Molecular Compounds)
81. NaHCO₃
Polyatomic, sodium hydrogen carbonate (or sodium bicarbonate)
82. Aluminum selenide
Binary, Al₂Se₃
83. N₂O
Molecular, dinitrogen monoxide
84. HClO
Ternary Acid, hypochlorous acid
85. O₂F
Molecular, dioxygen monofluoride
86. SBr₆
Molecular, sulfur hexabromide
87. Li₂Te
Binary, lithium telluride
88. Carbon monoxide
Molecular, CO
89. Carbon dioxide
Molecular, CO₂
90. P₂S₃
Molecular, diphosphorus trisulfide
91. Silver chloride
Binary, AgCl
92. Zinc nitrate
Polyatomic, ZnNO₃
93. Sodium nitride
Binary, Na₃N
94. Xenon hexafluoride
Molecular, XeF₆
95. I₂
Molecular, iodine (just iodine, see the POGIL on Molecular Compounds)
96. Ammonium nitrate
Polyatomic, NH₄NO₃
97. Tin(II) oxide
Multivalent, SnO
98. Sodium hydroxide
Polyatomic, NaOH
99. NO₂
Molecular, nitrogen dioxide
100. BF₃
Molecular, boron trifluoride