

# Science Notebook Layout **DON'T COPY UNDERLINED TEXT**

**Mrs. Aguirre's Webpage:** <http://www.quia.com/profiles/caguirre>

## Atoms: CPO pages 118-124 1/10/12

Definitions:

1. electron
2. Proton
3. Neutron
4. Nucleus
5. Atomic number
6. Isotope
7. Mass number

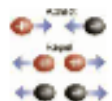
### 1. Three subatomic particles p 121 Fill in missing information

Particle	Where	Charge	Relative mass	Mass
Electron	Outside nucleus			0 (zero)
Proton	Nucleus			1 amu (atomic mass unit)
Neutron	Nucleus			1 amu

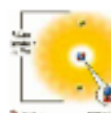
For boxes 2, 3, 4 below, copy the picture from the book. Also, write notes that describe what is shown in the picture.

Stamped up to 4 times

### 2. Charge on subatomic particles (pictures and text from pg 118)



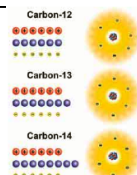
### 3. Atomic structure (how parts are put together) Pictures and text from pg. 121



### 4. Mass of subatomic particles p 121

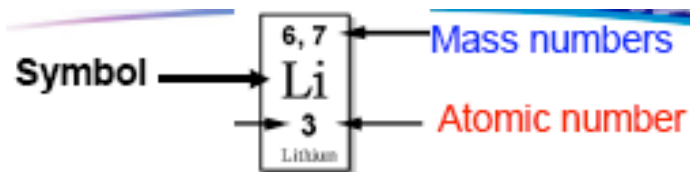


### 5. Isotopes of Carbon P 124



## What's Inside an Atom? 1/9/12

- BLUE = NEUTRONS
- RED = PROTONS
- YELLOW = ELECTRONS



A) # below symbol = atomic number

Atomic # = # of protons (reds)

B. # above the symbol = mass number

Mass number = particles inside the nucleus (reds and blues)

C. Sometimes there is more than one mass number- Atoms can have different numbers of neutrons (blues)

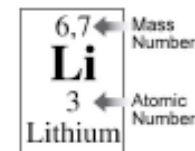
These are called isotopes.

D) What about yellows?

Yellow = electrons

Electrons equal protons

Yellow = Red



The Three Rules

- 1 # Red = Atomic Number
- 2 # Red = # Yellow
- 3 # Red + # Blue = Mass Number