

Rutherford Model	Hit the Penny Model
Alpha particle (+2 protons, 2 N)	
Nucleus of gold atoms	
Empty space of gold atom	
Electrons	

What does the "Hit the Penny" model mean?

Why did so few alpha particles hit the nucleus?

Why did they bounce off? (2 reasons)

Why did so many alpha particles miss the nucleus?
 Because the atom is mostly....



PROBLEM: What is the Structure of the Inside of an atom?

Watch video segment- write 3 notes

HYPOTHESIS:

Rutherford expected the positively charged alpha particles to pass through the gold atom

Person 1		Person 2	
Hits	Misses	Hits	Misses
<p>Leave 3 lines for tallying hits and misses</p>			

Calculate the percentage of hits by dividing:

$$\frac{\# \text{ hits}}{\# \text{ tries}} \times 100 = \underline{\quad} \%$$

