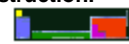


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HONORS WORKSHEET 3g: Nomenclature of acids



There is more than one way to define an acid, but in TOPIC 3 we will define it as a compound that is formed when an anion (a negative ion) has hydrogen ions (H^+) added to it so the charge becomes neutral.

There are two types of acid nomenclature that you will see in this TOPIC; binary acids formed when a simple, monatomic ion has hydrogen ions added and polyatomic (oxoacids) formed when polyatomic ions containing oxygen have hydrogen ions added.

1. Write the formula of each of the following acids. (14)

- (a) Nitric acid
- (b) Chloric acid
- (c) Hydrochloric acid
- (d) Sulfurous acid
- (e) Chlorous acid
- (f) Hydrobromic acid
- (g) Phosphoric acid
- (h) Nitrous acid
- (i) Perchloric acid
- (j) Hydrofluoric acid
- (k) Perbromic acid
- (l) Sulfuric acid
- (m) Bromic acid
- (n) Hypoiodous acid



2. Name the following acids. (14)

- (a) HClO_3
- (b) H_3PO_4
- (c) HI
- (d) H_2SO_3
- (e) HNO_3
- (f) HF
- (g) $\text{HC}_2\text{H}_3\text{O}_2$
- (h) HBr
- (i) H_3PO_3
- (j) HClO
- (k) H_2CO_3
- (l) H_2SO_4
- (m) HBrO_2
- (n) HNO_2