

# Models of Atoms throughout History

1808: Dalton Model  
-tiny solid spheres

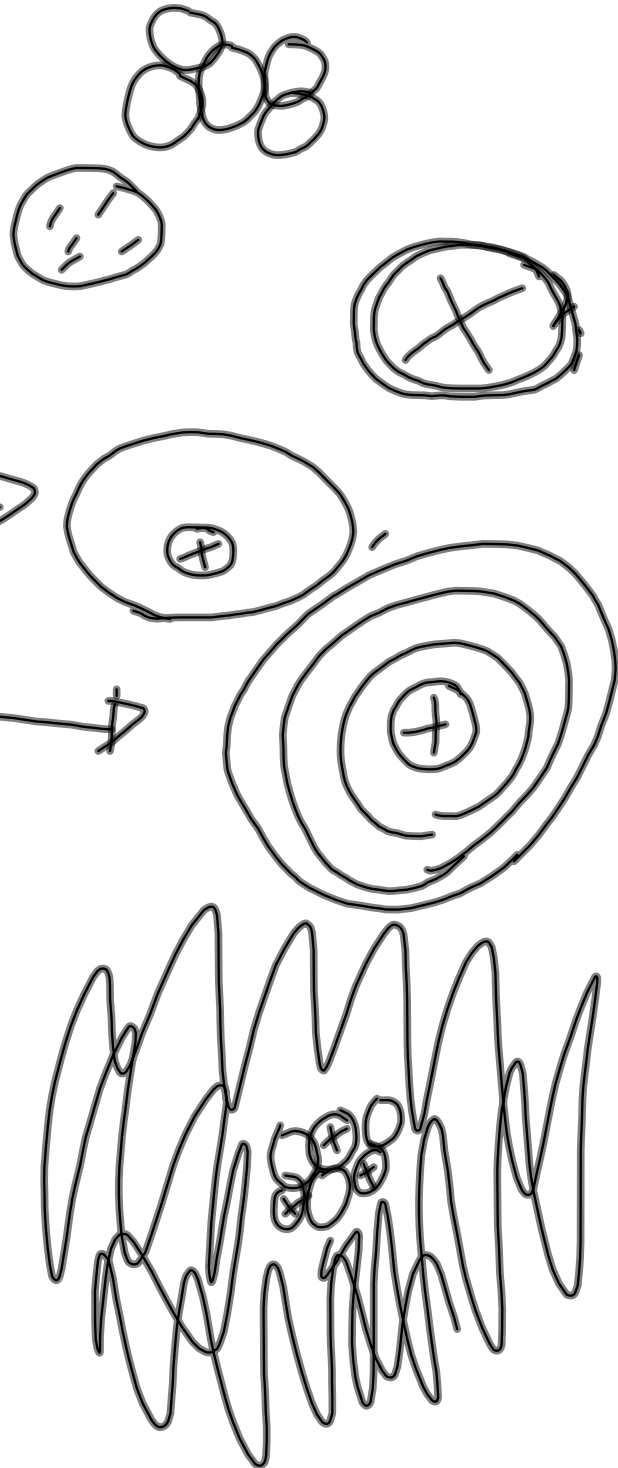
1897: Thomson Model  
-chocolate chip cookie

1904: Nagaoka Model  
-large positive center

1911: Rutherford Model  
- tiny nucleus  
<http://www.mhhe.com/physsci/chemistry/essentialchemistry/flash/ruther14.swf>

1913- Bohr Model  
- orbits

Today: The Modern Model  
- electron cloud



## Properties of Elements

<i>Atomic #</i>	6
<i>Symbol</i>	C Carbon
	12.01

<u>Natural Isotopes of Carbon</u>	
Mass Number	Natural Abundance
12	98.89%
13	1.11%

<http://education.jlab.org/itselemental/iso006.html>

<u>Property</u>	<u>Definition</u>
Symbol	latin abbreviation of the element's name
Atomic Number	The number of protons in the nucleus
Mass Number	the sum of the protons and neutrons in the nucleus
Atomic Mass	average mass number by percentage in natural existence

# Section 1 Assessment

## ELEMENT FACT SHEETS

1. Element Name
2. # of Protons
3. Classification on the Periodic Table
4. Atomic Mass #
5. Density
6. What are the most common isotopes?
7. How did the element get its name? Was it named after a person, country place or something else?
8. What is it used for?
9. Include a photograph of the element if possible. If not, include a model.

**MUST BE TYPED AND ORGANIZED IN A VISUALLY APPEALING WAY.**